

Calculating the number of sub-atomic particles in an atom

1. Complete the following table:

Particle	Position	Relative Charge	Relative Mass
Proton			
Neutron			
Electron			

2. In the table above, why do we want to know the relative charge and relative mass, and not the real charges in C (coulomb) and real masses in kg?

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3. Explain what mass number means.

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4. Explain what atomic number means.

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5. What are isotopes?

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6. Complete the following table.

The first two rows have been done for you. You may need to look at your periodic table.

Element	Symbol	Protons	Neutrons	Electrons
lithium	Li	3	$7-3=4$	3
carbon	C	6	$12-6=6$	6
sodium				
aluminium				
	Pb			
	Ti			
	Zn			
		72		
				15
			0	
tungsten				