

# Atomic structure – what do you recall?

## Atoms

Complete the sentences below.

Atoms have a small central .....

This is made up of protons and neutrons, with ..... around it.

## Subatomic particles

Complete the table below.

Name of subatomic particle	Relative mass	Relative charge
proton		
neutron	1	
		-1

## Numbers of subatomic particles

Complete the sentences below. You will not need to use all these words.

electrons	element	mass	nucleus
periodic	atomic	protons	no

In an atom, the number of electrons is equal to the number of .....

Atoms have ..... overall electrical charge.

All atoms of a particular ..... have the same number of protons.

Atoms of different elements have different numbers of ..... in their nucleus.

The number of protons in an atom is called its ..... number (proton number).

Atoms are arranged in the modern ..... table in order of their proton number.

The total number of protons and neutrons in an atom is called its ..... number.

## Isotopes

Describe what isotopes of an element are.

.....

.....

.....

## Atomic structure – what do you recall? – ANSWERS

### Atoms

nucleus, electrons

### Subatomic particles

Name of subatomic particle	Relative mass	Relative charge
proton	1	+1
neutron	1	0
electron	$\frac{1}{1836}$ (negligible)	-1

### Numbers of subatomic particles

electrons, no, element, protons, atomic, periodic, mass

### Isotopes

- atoms with the same number of protons (and electrons) but different numbers of neutrons
- atoms with the same atomic number but different mass numbers

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