

Formulae of ionic compounds – practice

Your task

Combine the positive and negative ions in the table below to write 15 correct formulae.

In a chemical formula you need to:

- have an equal number of positive charges and negative charges
- write the symbol for each ion **without** its charges
- write the symbol for a polyatomic ion inside brackets, **if** you need two or more of that ion
- write the number of each ion needed as a subscript to the right of its symbol.

Five formulae have been done for you. Make sure you understand why they are correct before starting.

	Cl ⁻	OH ⁻	NO ₃ ⁻	O ²⁻	SO ₄ ²⁻
K ⁺	KCl			K ₂ O	
NH ₄ ⁺		NH ₄ OH			
Mg ²⁺			Mg(NO ₃) ₂		
Al ³⁺					Al ₂ (SO ₄) ₃

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NH ₄ ⁺	NH ₄ Cl	NH ₄ OH	NH ₄ NO ₃	(NH ₄) ₂ O	(NH ₄) ₂ SO ₄
Mg ²⁺	MgCl ₂	Mg(OH) ₂	Mg(NO ₃) ₂	MgO	MgSO ₄
Al ³⁺	AlCl ₃	Al(OH) ₃	Al(NO ₃) ₃	Al ₂ O ₃	Al ₂ (SO ₄) ₃

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NH ₄ ⁺	NH ₄ Cl	NH ₄ OH	NH ₄ NO ₃	(NH ₄) ₂ O	(NH ₄) ₂ SO ₄
Mg ²⁺	MgCl ₂	Mg(OH) ₂	Mg(NO ₃) ₂	MgO	MgSO ₄
Al ³⁺	AlCl ₃	Al(OH) ₃	Al(NO ₃) ₃	Al ₂ O ₃	Al ₂ (SO ₄) ₃